

**3.3.3 Chemistry Paper 3 (233/3)**

**1** You are provided with:

- 1.60g of solid **A**, a dibasic acid.
- Solution **B** containing 4.75g per litre of salt **B**.
- Aqueous sodium hydroxide, solution **C**.
- Phenolphthalein indicator.

You are required to prepare a solution of solid **A** and use it to determine the:-

- Concentration of sodium hydroxide, solution **C**
- React salt **B** with excess sodium hydroxide and then determine the relative molecular mass of salt **B**.

**Procedure I**

- Using a burette, place  $25.0\text{cm}^3$  of solution **B** in each of two 250ml conical flasks. Using a pipette and **pipette filler**, add  $25.0\text{cm}^3$  of solution **C** to each of the two conical flasks. (The sodium hydroxide added is in excess). Label the conical flasks 1 and 2.
- Heat the contents of the first conical flask to boiling and then let the mixture boil for 5 minutes. Allow the mixture to cool.
- Repeat procedure (b) with the second conical flask.

While the mixtures are cooling, proceed with procedure II.

**Procedure II**

- Place **all** of solid **A** in a 250 ml volumetric flask. Add about  $150\text{cm}^3$  of distilled water, shake well to dissolve the solid and then add water to make up to the mark. Label this as solution **A**.
  - Place solution **A** in a clean burette. Using a pipette and **pipette filler**, place  $25.0\text{cm}^3$  of solution **C** in a 250ml conical flask. Add 2 drops of phenolphthalein indicator and titrate with solution **A**. Record your results in Table 1. Repeat the titration two more times and complete the table.
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Table 1

	I	II	III
Final burette reading			
Initial burette reading			
Volume of solution A used (cm <sup>3</sup> )			

(4 marks)

Calculate the:-

- (i) average volume of solution A used: (½ mark)
- (ii) concentration in moles per litre of the dibasic acid in solution A; (2 marks)  
(Relative molecular mass of A is 126).
- (iii) moles of the dibasic acid used; (1 mark)
- (iv) moles of sodium hydroxide in 25.0cm<sup>3</sup> of solution C. (1 mark)
- (v) concentration of sodium hydroxide in moles per litre. (2 marks)

### Procedure III

Add 2 drops of phenolphthalein indicator to the contents of the first conical flask prepared in procedure I and titrate with solution A. Record your results in Table 2. Repeat the procedure with the contents of the second conical flask and complete the table.

Table 2

	1st conical flask	2nd conical flask
Final burette reading		
Initial burette reading		
Volume of solution A used (cm <sup>3</sup> )		

(3 marks)

Calculate the:-

- (i) average volume of solution A used; (½ mark)
- (ii) moles of the dibasic acid used; (1 mark)
- (iii) moles of sodium hydroxide that reacted with the dibasic acid. (1 mark)

- (iv) moles of sodium hydroxide that reacted with  $25.0\text{cm}^3$  of salt **B** in solution **B**; (2 marks)
- (v) Given that 1 mole of salt **B** reacts with 2 moles of sodium hydroxide, calculate the:
- I. number of moles of salt **B** in  $25.0\text{cm}^3$  of solution **B**; (1 mark)
  - II. concentration in moles per litre of salt **B** in solution **B**; (1 mark)
  - III. relative molecular mass of salt **B**; (2 marks)

- 2 (a) You are provided with solid **D**. Carry out the following tests and write your observations and inferences in the spaces provided.
- (i) Place about one half of solid **D** in a test-tube and heat it strongly. Test any gas produced with both red and blue litmus papers.

Observations (2 marks)	Inferences (1 mark)
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- (ii) Place the rest of solid **D** in a boiling tube. Add about  $10\text{cm}^3$  of distilled water. Shake well.

To a  $2\text{cm}^3$  portion of the solution, add about  $1\text{cm}^3$  of hydrogen peroxide and shake well. To the resulting mixture, add aqueous sodium hydroxide dropwise until in excess.

Observations (1 mark)	Inferences (1 mark)
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- (b) You are provided with solution **E**. Carry out the following tests and write your observations and inferences in the spaces provided. Divide solution **E** into **two** portions.
- (i) To one portion of solution **E** in a test-tube, add 3 drops of barium nitrate. **Retain the mixture for use in test (ii) below.**

<b>Observations</b>	<b>Inferences</b>
(1 mark)	(2 marks)

- (ii) To the mixture obtained in (i) above, add about 5 cm<sup>3</sup> of 2M nitric (V) acid.

<b>Observations</b>	<b>Inferences</b>
(1 mark)	(1 mark)

- (iii) To portion two of solution E in a test-tube, add 2 drops of acidified potassium dichromate (VI) and warm the mixture.

- 3 You are provided with liquid F. Carry out the following tests and record your observations and inferences in the spaces provided.

- (a) Place five drops of liquid F on a clean dry watch glass and ignite it.

<b>Observations</b>	<b>Inferences</b>
(1 mark)	(1 mark)

- (b) Place about 2cm<sup>3</sup> of liquid F in a clean dry test-tube, add all the sodium hydrogen carbonate provided.

<b>Observations</b>	<b>Inferences</b>
(1 mark)	(1 mark)

- (c) Place about 2cm<sup>3</sup> of liquid F in a test-tube, add about 1cm<sup>3</sup> of acidified potassium dichromate (VI) and warm the mixture.

<b>Observations</b>	<b>Inferences</b>
(1 mark)	(1 mark)